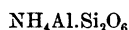


Dr. F. W. Clarke and was entitled, 'The Action of Ammonium Chlorid upon certain Silicates,' by F. W. Clarke and George Steiger.

The authors described a series of experiments in which various silicates were heated in a sealed tube to 350° C. with dry ammonium chlorid. After leaching out the contents of the tube with water it was found that alkalis were removed as chlorids and replaced by ammonia, analcite and leucite are thus transformed into an ammonium leucite:



which is perfectly stable at 300° and only begins to decompose when heated in the open air to 350°.

Some eight other silicates were given preliminary study and the reaction was found to be fairly general. The product from natrolite contained 8.3 per cent. of ammonia and other zeolites took up from four to six per cent. The investigation is to be continued.

The fourth paper was read by Dr. F. K. Cameron and was entitled, 'Hydrochloric Acid and Aqueous Phenol,' by F. K. Cameron and J. A. Emory.

The authors determined the freezing-point curve for hydrochloric acid solutions, saturated with respect to phenol. Each independently determined the concentrations of the various solutions and their freezing-points for inter-comparison. The curve was found to be a straight line, parallel to the curve for water and hydrochloric acid alone, from which it would seem that the solubility of phenol is practically constant through the range of temperature involved, and the lowering of the freezing-point of the solvent is a purely additive effect of the two solutes.

The fifth paper was read by Dr. F. K. Cameron and was entitled, 'The System Water, Hydrochloric Acid and Phenol,' by F. K. Cameron and W. H. Krug.

On lowering the temperature of the system, solid phenol separates. But if the initial mass of water be relatively large its concentration with respect to hydrochloric acid is practically unaffected, while the solid phenol is separating and consequently the temperature of the phenol remains very constant. The freezing-point

curve for phenol in contact with aqueous solutions of hydrochloric acid of various concentrations was determined. Its practical value for a rapid determination of the approximate strength of hydrochloric acid solutions was indicated.

WILLIAM H. KRUG,
Secretary.

NEW YORK SECTION OF THE AMERICAN
CHEMICAL SOCIETY.

THE regular meeting of the New York Section of the American Chemical Society was held on Friday evening, the 5th inst., at the Chemists' Club, and was well attended, over sixty members and their friends being present. Dr. C. F. McKenna occupied the chair, calling the meeting to order at 8:30 p. m.

After electing four delegates to represent the Section in the Council, the following papers were read:

(1) 'The Importance and Trend of Recent Work on the Chemistry of Life and the Products of Life,' by Jerome Alexander.

(2) 'A Preliminary Study of the Cobalt-Cyanides,' by E. H. Miller and J. A. Mathews.

(3) 'The Chemistry of Corn Oil. First Paper: Determination of the Constants,' by Herman T. Vulté and Harriet W. Gibson.

(4) 'A Practical Electric Furnace,' by A. J. Rossi.

Mr. Rossi exhibited a practical and easily constructed electric furnace with which he has prepared some very rich Titanium alloys, a specimen of which was exhibited with an invitation to break off pieces as samples. Although a sledge hammer was supplied no samples were taken. Arrangements are progressing toward the preparation of these alloys on a large scale for the steel trade.

DURAND WOODMAN,
Secretary.

TORREY BOTANICAL CLUB.

AT the meeting on November 29th, the scientific program consisted of a paper by Dr. C. C. Curtis, on Seaweeds, with lantern views illustrating the chief families and with a condensed summary of the modes of reproduction and other characteristics of each. Dr. Curtis also gave brief directions respecting methods of collecting and preserving the marine algæ, urging the collector to make microscopic study of all