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New chloride of manganese

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It appears therefore that hordein differs very materially from this substance as well as from saw-dust, of which the analysis is thus given by MM. Gay-Lussac and Thenard :

Carbon	52·
Oxygen	42·4
Hydrogen	5·6
	<hr style="width: 10%; margin: 0 auto;"/>
	100·

M. Marcet therefore considers hordein a peculiar substance, but most resembling starch. The azote which it contains may, he thinks, be derived from the presence of gluten.

The following are the results of M. Marcet's analysis of gluten and yeast :

	Gluten.	Yeast.
Carbon	55·7	30·5
Oxygen	22·0	57·4
Hydrogen	7·8	4·5
Azote	14·5	7·6
	<hr style="width: 10%; margin: 0 auto;"/>	<hr style="width: 10%; margin: 0 auto;"/>
	100·	100·

Ibid. p. 27.

NEW CHLORIDE OF MANGANESE.

M. Dumas obtained this compound by putting a solution of manganesic acid into contact with sulphuric acid and fused common salt. Water and the new chloride are formed; the former is retained by the acid, the latter volatilizes in a gaseous form, with a greenish tint, and when passed into a tube cooled to 5° or 4° of Fahrenheit, it condenses into a liquid of a brownish-green colour. The most simple process appears to be to form a common green chameleon, to convert it into red chameleon by sulphuric acid, and to evaporate the solution, which will give a residue consisting of sulphate and manganesate of potash. This mixture acted upon by concentrated sulphuric acid produces the solution of manganesic acid, into which the common salt is to be thrown in small pieces until the vapours which rise are colourless; the latter effect is a sign that all the manganesic acid is decomposed, and that muriatic acid only is produced. This chloride of manganese corresponds in proportions to the manganesic acid; it is readily formed and examined, but not easily preserved. An analogous compound is obtained when a fluoride is used instead of the chloride, but a sufficient quantity for examination has not yet been procurable.—*Ibid.* xxxvi. 81.

ON THE POWER OF WATER AND BROMINE IN CONDUCTING ELECTRICITY.

M. de la Rive found, as had been previously ascertained by M. Balard, that pure dry bromine did not conduct the electricity of a voltaic battery, consisting of sixty pairs of plates very strongly charged, a delicate galvanometer being the test. A similar experiment was then made with pure water contained in a glass capsule, and communicated